

ONE-STEP SYNTHESIS TO ENHANCE THE ACIDITY OF A BIOCARBON-BASED SULFONATED SOLID ACID CATALYST

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ABSTRACT

The main purpose of this study is to produce and generate a solid acid catalyst from biomass with high reactivity that can be used in catalytical reactions such as hydrolysis, and is environmentally friendly and reusable. A biocarbon-based sulfonated catalyst was prepared by the carbonization of palm empty fruit bunches (PEFB), followed by sulfonation. In order to enhance the acidity of the biocarbon, different concentrations of hydroxyethylsulfonic acid were added to the solution during sulfonation at 180°C for 4 h in a Teflon stainless steel autoclave. The H⁺ ion capacity of the biocarbon-sulfonated acid catalyst (BSC) was increased twofold (3.57 mmol/g) in the presence of 10% of hydroxyethylsulfonic acid and 10% of acrylic acid. X-Ray Fluorescence (XRF) analysis showed that the BC-SO₃H contained 38% of S. The original structure of the PEFB after carbonization disintegrated from the fibrous materials onto porous carbon. The crystalline index (CrI) of the PEFB significantly decreased to about 32% and a wide broad peak of a X-Ray Diffraction (XRD) pattern of around 20–30° were observed, which shows that an amorphous biocarbon structure had been identified. Fourier Transform Infra-Red (FT-IR) analysis confirmed that the -SO₃H, COOH and -OH functional groups were deposited on the carbon due to specific peaks at around 1180 cm⁻¹, 1724 cm⁻¹ and 3431 cm⁻¹, respectively. Decomposition of the sulfonic groups on the biocarbon-sulfonated solid catalyst was observed from 227.9°C, as it shown by thermal gravimetric analysis (TGA).

Keywords: Acid catalyst; Biocarbon; Palm empty fruit bunch; Sulfonated; Sulfonation

1. INTRODUCTION

Palm is one of the most important commodities in Indonesia due to its rapid development. The major product from the palm industry is Crude Palm Oil (CPO), but with its increasing production, the waste, that takes the form of empty fruit bunches, has increased. Nowadays, biomass and industrial waste have become very interesting issues as aspects of catalyst development, both in research and from the technical point of view, due to their valuable merit of industrial waste (Guerrero-Pérez et al., 2006; Kusriani et al., 2018). Biomass energy is an ideal clean and renewable energy source, characterized by its wide range of sources, low prices, strong reproducibility and less pollution creation (Wenjing et al., 2018).

In the using of solid acid catalysts, they are easy and efficient when separated from their products, are reusable and it is possible to apply them in wide of applications, but most such catalysts developed are expensive and quite difficult to prepare (Okuhara, 2002).

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Recently, work on sulfonated solid acid catalysts has attracted great attention from researchers for the hydrolysis reaction of cornstarch (Nata et al., 2015; Nata et al., 2017b) and biodiesel production from waste cooking oil (Zong et al., 2007; Nata et al., 2017a). Performance in the reaction of carbon-derived catalysts is dependent on the precursor as raw materials for carbon production and treatment processes (Tao et al., 2015). From the point of view of “green chemistry”, the sulfonated carbon catalyst has emerged as a promising solid acid catalyst (Jiang et al., 2012).

Theoretically, at low carbonization (400–600°C), biomass generates a highly cross-linked, multi-ringed, aromatic structure anchored to lignin that can be easily functionalized with catalytically active acidic groups by slow pyrolysis (Kastner et al., 2012). Generally, a two-step process is involved in the production of sulfonated carbonaceous materials. Saccharide is incompletely carbonized at a temperature of > 400°C for >15 h under an inert atmosphere. A large amount of sulphuric acid use in the sulfonation process at a high temperature for the inactive surface of carbonaceous material (Zong et al., 2007). This process uses hazardous material and a large amount of harmful waste is produced; moreover, the carbon in the concentrate sulphuric also needs special attention for its separation and treatment.

Hydrothermal carbonization (HTC) is a thermochemical process capable of converting wet biomass into a carbon-enriched solid as hydrochar. The HTC process consists of several reactions conducted both in series and in parallel, including hydrolysis, dehydration, decarboxylation, condensation and aromatization (Merzari et al., 2018). HTC is process which involves the decomposition of several carbohydrates in aqueous solution at 180°C. This method is cheap, mild and environmental friendly, as no organic solvents, catalysts or surfactants are used (Titirici et al., 2007). In a previous study, Xiao et al. (2010) performed hydrothermal treatment with hydroxyethylsulfonic acid as a sulfonate agent to produce carbon from glucose and used it for an esterification process in order to examine its catalytic ability. However, this procedure only achieved 1.7 mmol/g of acidity and still owned little of functional groups. Therefore, to generate carbonaceous material loaded with carboxylic groups, known as an active group that participates in the reaction, acrylic acid was added (Bautista-Toledo et al., 2005). In order to produce a high content of functional groups on the carbon material, it is possible to modify the surface by a one-step HTC process for sulfonation and thus improve the acidity of the carbon.

This work focuses on the effect of hydroxyethylsulfonic acid concentration and the addition of acrylic acid during the hydrothermal process. Therefore, the characterization of aspects such as acidity, morphological structure, crystalline structure, functional groups and thermal gravimetric analysis was investigated.

2. METHODS

2.1. Materials

The raw material as a carbon source for the catalyst is palm empty fruit bunches, collected from PT. Perkebunan Nusantara XIII, South Kalimantan, Indonesia. Citric acid ($C_6H_5Na_3O_7 \cdot 2H_2O$), acrylic acid ($C_3H_4O_2$), hydroxyethylsulfonic acid ($C_2H_6O_4S$), sodium chloride (NaCl), sodium hydroxide (NaOH), oxalic acid ($H_2C_2O_4$), chloride acid (HCl), methanol (CH_3OH), and phenolphthalein were used in analytical standard and purchased from Sigma-Aldrich, St. Louis, USA.

2.2. Synthesis of a Biocarbon-sulfonated Solid Acid Catalyst

A biocarbon-sulfonated solid acid catalyst was conducted in line with Xiao et al. (2010) through one-step hydrothermal treatment with some modification. Empty fruit bunches were heated in a furnace at 350°C for 1 h, producing carbon meshed up to 60 mesh. The obtained

carbon material was mixed with hydroxyethylsulfoic acid at concentrations of 10, 20 and 30% v/v, 1.5 g of citric acid, 1.5 mL of acrylic acid and 30 mL of deionized (DI) water poured into a 50 mL Teflon-lined stainless steel autoclave, then heated for 4 h at 180°C. A black carbon material was produced, which was then washed with DI water and 50% of methanol twice. The biocarbon-sulfonated solid acid catalyst (BSC) was obtained after drying at 80°C for 6 h in an oven. The four different types of carbon particle produced were biocarbon without sulfonation, and carbon with sulfonation containing 10%, 20% and 30% of hydroxyethylsulfonic acid, namely BC, and BSC-1, BSC-2, and BSC-3, respectively.

2.3. Calculation of Acid Content

The Brønsted acid content in the sample was measured by neutralization titration (Onda et al., 2008). Carbon material (50 mg) was added into a 2 M NaCl (25 mL) solution at room temperature and stirred for 5 h. It was then filtered and a clear solution of supernatant was obtained; phenolphthalein was added and then titrated with 0.1 M NaOH to determine the acid content.

2.4. Characterization

Observation of the morphological structure image of the biocarbon was obtained by a scanning electron microscope (SEM) using JEOL, JSM-6500 LV equipment. The functional groups on the surface of the material were detected by Fourier transform infrared spectrometry (FT-IR, Bio-rad, Digilab FTS-3500) at a wavelength range of 4000-400 cm^{-1} and a scan rate of 8. X-ray diffraction (XRD) analysis was conducted on an X-ray diffractometer using Cu $K\alpha$ radiation (Rigaku D/MAX-B) with 2θ (10–40°) at a scan rate of 2° min^{-1} and with voltage and current of 40 kV and 100 mA, respectively. Thermal gravimetric analysis (TGA) performance using a Perkin Elmer, Diamond TG/DTA was conducted from 30°C-600°C in a nitrogen atmosphere at a heating rate of 10°C/min. The remaining material content in the sample was calculated by mass change from the corresponding TGA curves.

3. RESULTS AND DISCUSSION

The PEFB in its original form turned to black carbon after carbonization, which was followed by sulfonation via hydrothermal treatment. Hydrothermal carbonization is a method which is an attractive mild preparation of carbons with special properties. The accessible external surface is densely functionalized with polar groups, such as carboxylic and hydroxyl, and confers them high hydrophilicity and susceptibility to chemical modification (Fraile et al., 2012). The treatment of carbon materials under HTC increases or changes their solubility; melts any crystalline parts; accelerates the physical and chemical interaction between reagents and the solvents, facilitating ionic and acid reactions; and leads to carbonaceous structure formation (Hu et al., 2010). A schematic diagram of the preparation of a biocarbon-sulfonated solid acid catalyst is shown in Figure 1.

It is known that during the carbonization process, lignin will break into partial decompose, then hemicellulose and cellulose would decompose through thermal homolysis, hydrolysis, dehydration, and rearrangement of molecular reactions to form a polymerized aromatic structure, then next process is bounding sulfonic groups to biocarbon materials during hydrothermal treatment. The arrangement of biocarbon material by dehydration in the first step, then allows sulfonic and carboxylic acid groups will be induced on carbon materials, subsequently providing hydroxyl functional groups for intermolecular dehydration. In this case, hydroxyethylsulfonic acid was used as a functional molecule to introduce sulfonic groups to the biocarbon material, and carbonyl groups were supplied by citric acid, which helped obtain carbon with high acidity (Xiao et al., 2010).

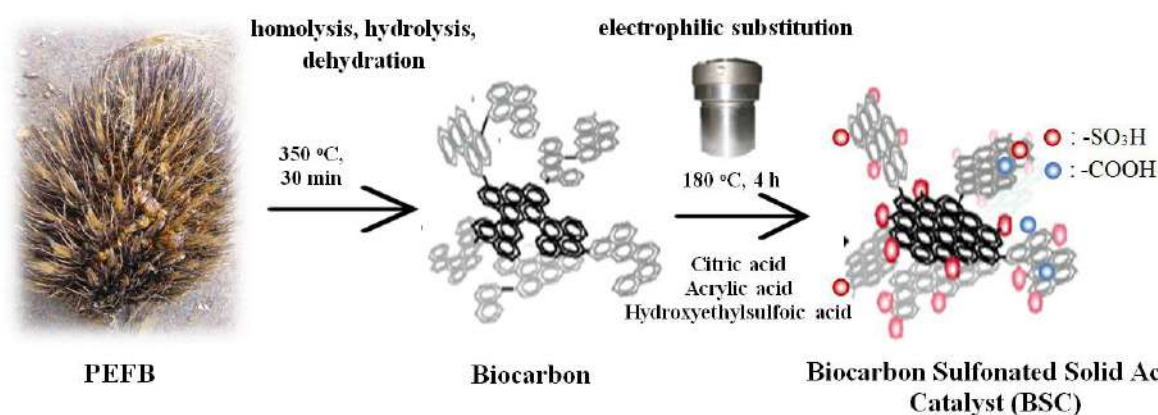


Figure 1 Schematic diagram of the preparation of a biocarbon-sulfonated solid acid catalyst from PEFB

Moreover, the addition of acrylic acid as a co-monomer in hydrothermal treatment causes the escalation of catalyst acidity. The acrylic acid induction for the hydrothermal carbonization process could be achieved for product with rich of carboxylic groups on the surface, which was known that carboxylic groups are more active than sulfonic groups and participate in the reaction instead of sulfonic groups (Demir-Cakan et al., 2009). The presence of sulfonic groups is reached by the reaction of the sulfonating agent and the polycyclic aromatic rings of biocarbon that make up the new structure, through electrophilic substitution (Aldana-Pérez et al., 2012; Lokman et al., 2016).

Table 1 shows the acidity and S content based on X-Ray fluorescence for different types of biocarbon-sulfonated solid acid catalyst. The addition of 10% of hydroxyethylsulfoic acid increased the level of S content by around 38%; however, the biocarbon acidity decreased when >10% hydroxyethylsulfoic acid was added. This was probably due to the competitiveness of each functional group when binding with biocarbon during hydrothermal treatment. The more the hydroxyethylsulfoic acid used in the sulfonation process, the higher the S content accommodated on the carbonaceous materials, resulting in more acidic groups such as carboxyl, carbonyl and hydroxyl not being able to bind with the biocarbon, whereas these acid groups acted as suppliers of H⁺ ions, which helped produce biocarbon with high acidity. This explains the decrease in acid capacity for BSC-2 and BSC-3 compared to BSC-1. A higher acidity content in the solid acid catalyst will improve its reactivity during reactions such as hydrolysis and esterification. In the case of the hydrolysis reaction, the H⁺ ions in the catalyst will break the polymer bond of amylose.

Table 1 Acidity and S content of different types of biocarbon solid acid catalyst

Sample	Acidity (mmol/g)	S content (%)
BC	0.209	-
BSC-1	3.569	38.0
BSC-2	2.380	46.8
BSC-3	2.712	54.4

Biocarbon morphological characterization of the solid acid catalyst was conducted in order to predict the physical and chemical properties of the material. As shown in Figure 2a, the PEFB had a fibrous shape with a smooth surface, while Figure 2b shows the structure after carbonization. In comparison, Figures 2c, 2d and 2e show the difference in surface morphology of biocarbon, the sulfonation process in the presence of hydroxyethylsulfoic and acrylic acid

were generated scattered shape form, but more aggregated. Presumably, hydroxyethylsulfonic acid acted as an agent that disintegrated the fibrous materials onto the porous carbon, which allowed the sulfonate groups to attach to the surface of the carbon. In addition, more particles were formed with a smaller size during dehydration, and polymerization caused the surface of the material to be no longer smooth due to the change in the material morphology induced by the addition of hydroethylsulfonic and acrylic acid.

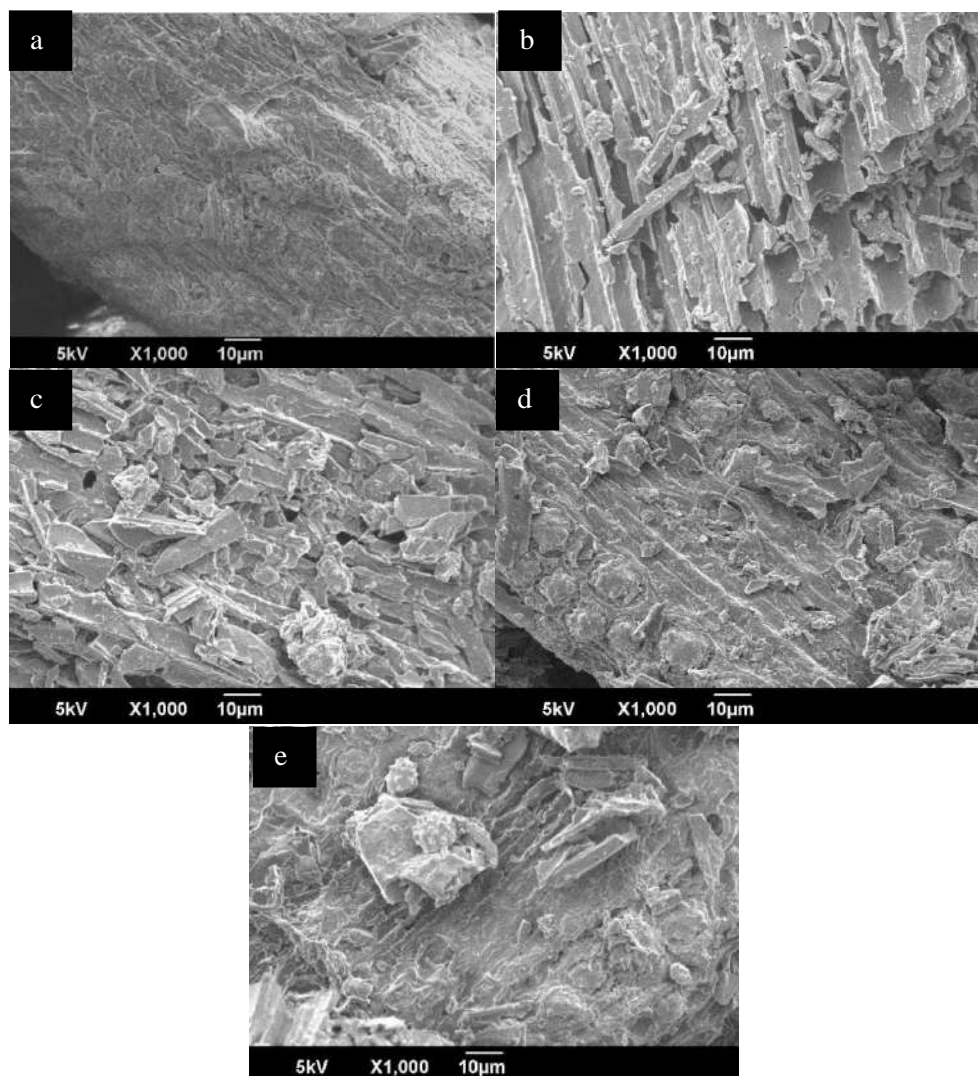


Figure 2 SEM images of: (a) PEFB; (b) BC; (c) BSC-1; (d) BSC-2; (e) BSC-3

Figure 3 shows the XRD pattern for the material whose crystalline PEFB index significantly decreased by around 22.38%. The BC, BSC-1, BSC-2 and BSC-3 formed a wide-broad peak pattern of around $2\theta = 20\text{--}30^\circ$, which shows the amorphous structure of carbon (Okamura et al., 2006). These data prove that treatment at different concentrations of acrylic acid does not effects to the amorphous structure of the carbon. The reduction in the crystalline index (CrI) indicates that the crystalline structure transformed into carbon due to the carbonization process which involved degradation of lignin, hemicellulose and cellulose.

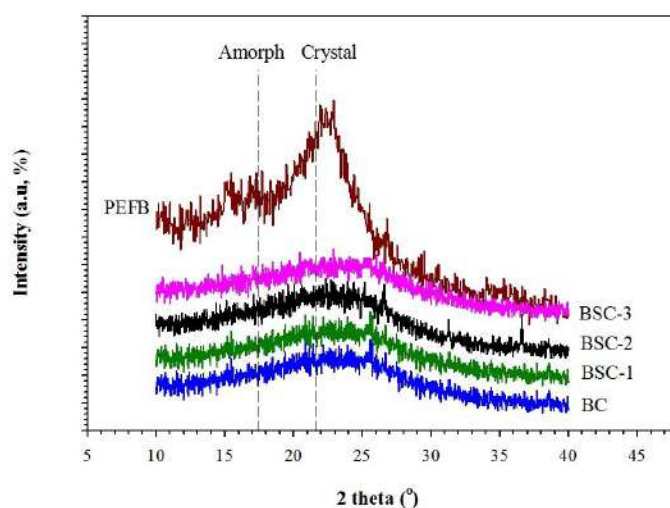


Figure 3 XRD patterns of PEFB, BC, BCS-1, BCS-2, and BCS-3

FT-IR measurement was made to identify the functional groups on the material. Absorption bands of -OH stretching at 3450 cm^{-1} , -OH bending at $1000\text{--}1300\text{ cm}^{-1}$ and C=C double bonds at 1608 cm^{-1} were found in all the samples (Zong et al., 2007). The presence of a sulfonate absorption band at 1207 cm^{-1} and C=O double bonds at 1720 cm^{-1} , which indicate carboxylic groups, was identified for BSC-1, BSC-2, BSC-3, as shown in Figure 4. The different concentrations of hydroxyethylsulfonic acid give different intensity peaks for the sulfonate group, a higher concentration, and sharper peaks were formed. This observation is related to the XRF results. Furthermore, FT-IR analysis proved that the biocarbon materials were successfully sulfonated during the hydrothermal process with the addition of hydroxyethylsulfonic acid. The presence of sulfonic acid groups (-SO₃H) on the carbon will change the pH of the surface, therefore they also change the adsorption characteristics and makes these carbons suitable for the application as heterogeneous solid acid catalysts (Fraga et al., 2016).

In order to understand the typical degradation of PEFB and the impact of sulfonation on the thermal stability of biocarbon-sulfonated solid acid catalysts, thermal gravimetric analysis was conducted. In Figure 5, the gravimetric analysis of all the materials carbonized up to 600°C is shown. The PEFB starts the dehydration of water at 74.4°C , followed by the decomposition of lignin and hemicellulose at 332.5°C . Major decomposition of the biocarbon material, including the sulfonated solid catalyst, was observed from 227.9°C for the sulfonic group.

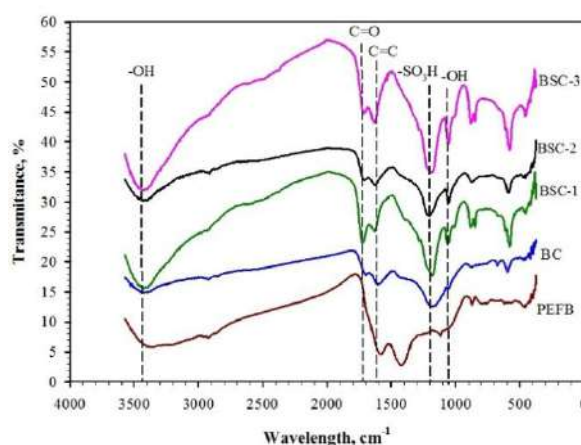


Figure 4 FT-IR spectra of PEFB, BC, BSC-1, BSC-2, and BSC-3

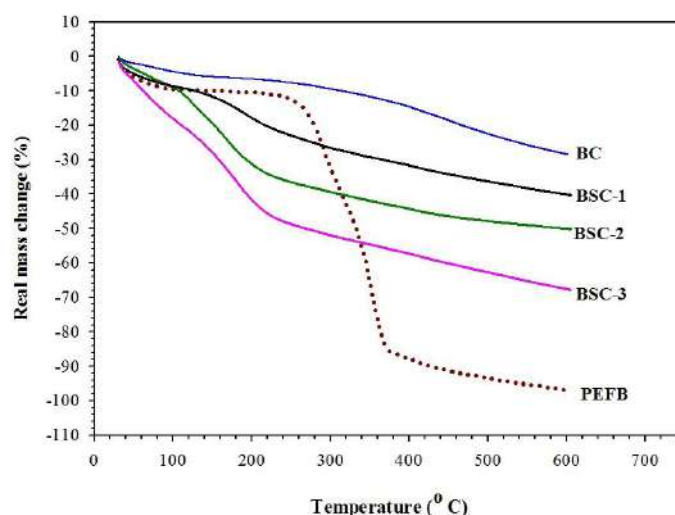


Figure 5 TGA curves of PEFB, BC, BSC-1, BSC-2, BSC-3

Sulfonation of the biocarbon lowered the temperature at which mass loss occurred and partially oxidized the structure, reducing thermal stability (Kastner et al., 2012). A higher concentration of sulfonic acid will result in more mass loss on the TGA curve, as for BSC-1, BSC-2 and BSC-3 at about 40%, 50% and 67%, respectively. This phenomenon of mass loss indicates that the material contained in the catalyst rapidly diminishes as the temperature increases, and is related to the number of sulfonic groups obtained by the XRF results. It can be concluded that a biocarbon-based solid catalyst sulfonated with a higher amount of sulfonate agent will have a frailer structure and be unstable with alterations in temperature compared to the original biocarbon or the biocarbon catalyst sulfonated with less sulfonate agent.

4. CONCLUSION

The strong acid content, rich of sulfonic and carboxylic groups of materials could be easily synthesized by a one-step hydrothermal process using biocarbon from incomplete carbonization of PEFB, and in the presence of hydroxyethylsulfonic, acrylic and citric acid in mild conditions. The simplicity of operation, high activity and stability, low cost of raw materials and reusability are the main features of this original biocarbon-based sulfonated solid acid catalyst, which demonstrates that biocarbon has great potential for green processes in various catalytic applications.

5. ACKNOWLEDGEMENT

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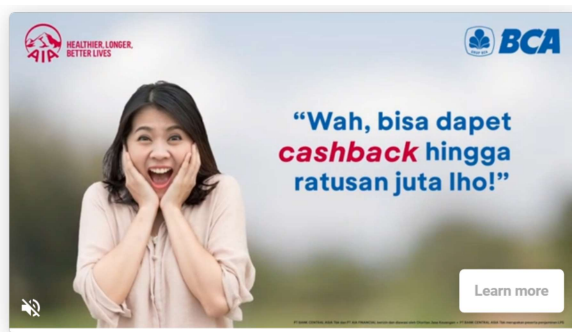
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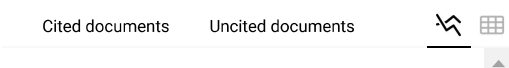
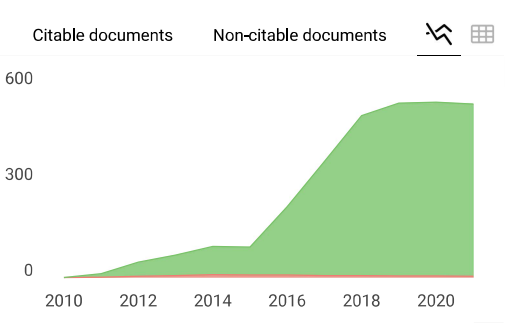
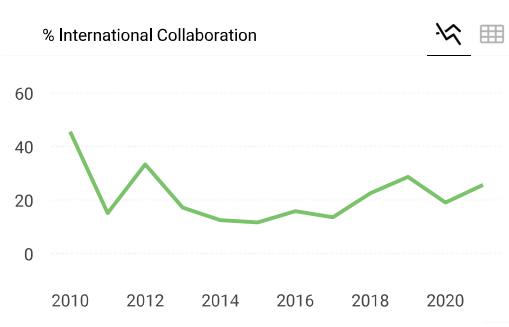
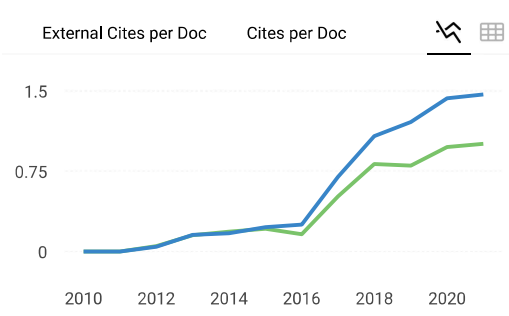
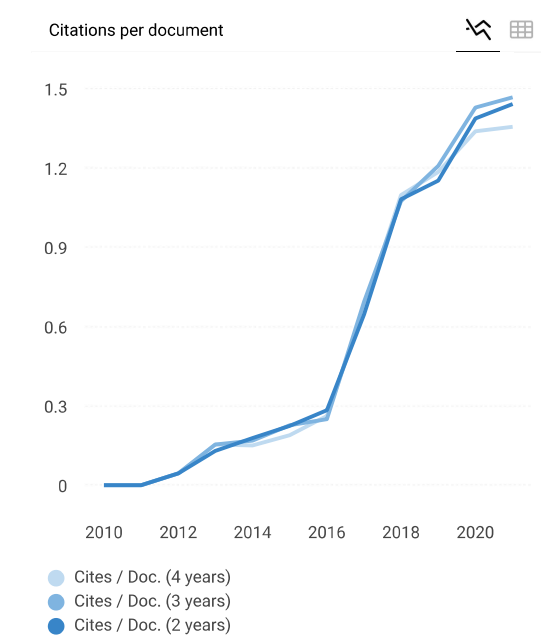
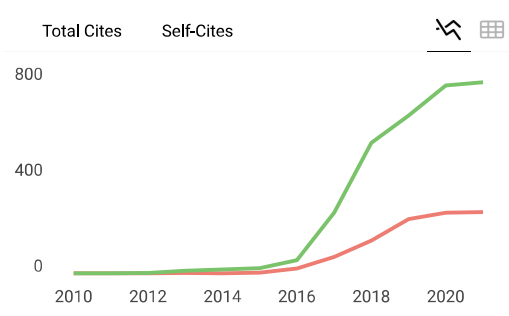
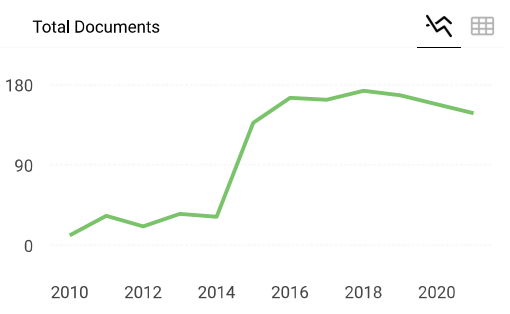
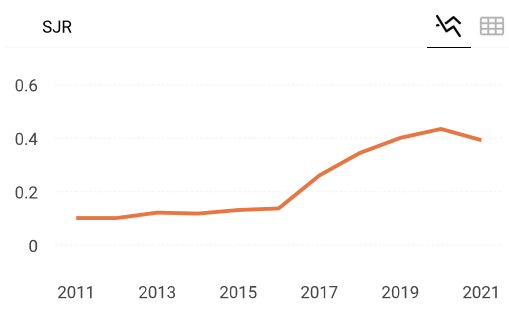
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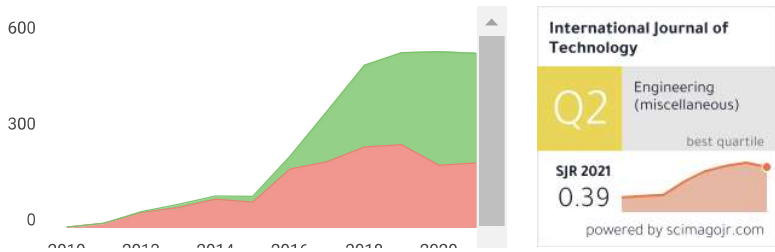
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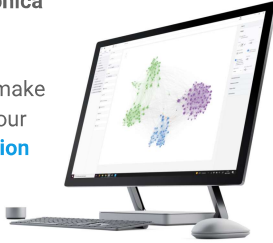


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We are sorry to tell you that SCImago Journal & Country Rank is not a journal. SJR is a portal with scientometric indicators of journals indexed in Elsevier/Scopus.

We suggest you visit the journal's homepage (See submission/author guidelines) or contact the journal's editorial staff, so they could inform you more deeply.

Best Regards, SCImago Team

A **anju panwar** 2 years ago

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